

**Supporting Information for**

**Optimizing the process parameters for preparation of powdered activated coke (PAC) to achieve maximum SO<sub>2</sub> adsorption capacity using response surface methodology (RSM)**

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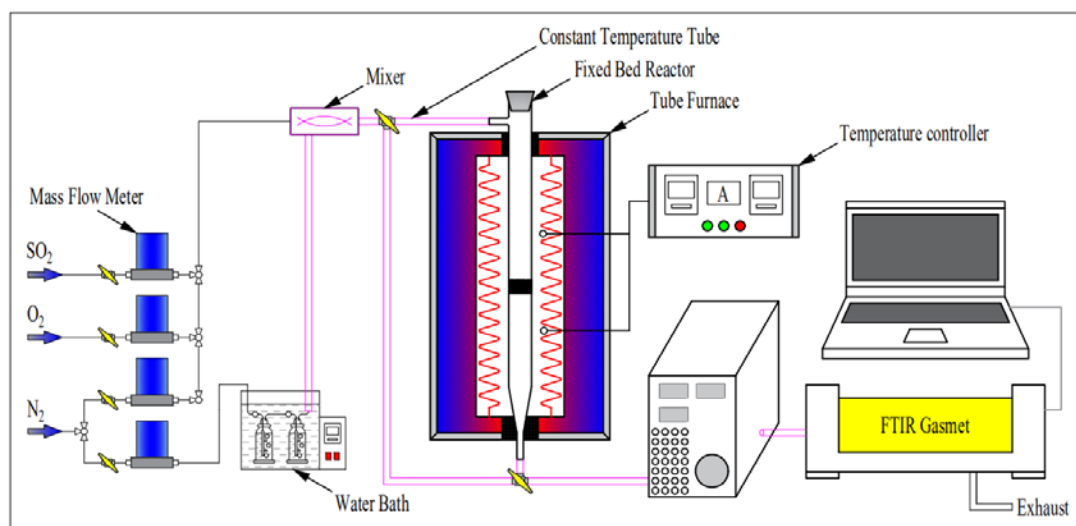
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### **Fixed-bed SO<sub>2</sub> adsorption system and the adsorption condition:**

The SO<sub>2</sub> adsorption capacity of the PAC was evaluated at 75°C using a fixed bed experimental system, as shown in Fig. S1. This system mainly consists of a fixed-bed glass reactor (inner diameter is 17 mm, wall thickness is 1.5 mm), an electric heater, a hot water bath, a portable Fourier transform infrared (FTIR) instrument from Finland Gasmeter Company, valves and mass flow meters. Specifically, 1g of sample was loaded into the glass reactor, when the temperature of the reaction zone reached the required value and maintained a stable state, the total flow rate (0.5 L/min) was introduced into the reactor by simulating flue gas (1500 ppm of SO<sub>2</sub>, with 6% O<sub>2</sub>, with 8% H<sub>2</sub>O, N<sub>2</sub> balance). The H<sub>2</sub>O stream was introduced by a hot water bath and controlled by bubbling a certain amount of N<sub>2</sub>. We continuously monitored the concentrations of SO<sub>2</sub>/H<sub>2</sub>O in the inlet and outlet of the fixed-bed reactor using the FTIR. Then we used the following formula to calculate the SO<sub>2</sub> adsorption capacity ( $S$ ) of the experimental samples.

$$S=28.6 \times \int_0^t \frac{(C_{in}-C_{out})V}{M} dt, (S1)$$

where,  $C_{in}$  and  $C_{out}$  are the concentrations of SO<sub>2</sub> (ppm) of the reactor inlet and outlet;  $V$  is the total flow rate (L/min) of simulated flue gas;  $M$  is the weight (g) of the sample packed into the glass reactor.



**Fig. S1.** Fixed-bed SO<sub>2</sub> adsorption system.

### **Pore structural characterization:**

N<sub>2</sub> adsorption-desorption isotherm of PAC-OP, Commercial AC1 and Commercial AC2 were obtained at 77 K by a Micromeritics ASAP 2020 analyzer. The surface area ( $S_{\text{BET}}$ ) was estimated using the Brunauer-Emmett-Teller (BET) multi-points method. The micropore surface ( $S_{\text{mic}}$ ) was calculated by the t-plot method. The micropore volume ( $V_{\text{mic}}$ ) was obtained by the DR method. The total pore volume ( $V_{\text{tot}}$ ) was calculated by the analysis software at relative pressure larger than 0.97 [S1, S2]. The pore size distributions of each sample were analyzed with the QSDFT method, and through which the ultra-micropore volume ( $V_{\text{mic} < 1 \text{ nm}}$ ) could be obtained [S3].

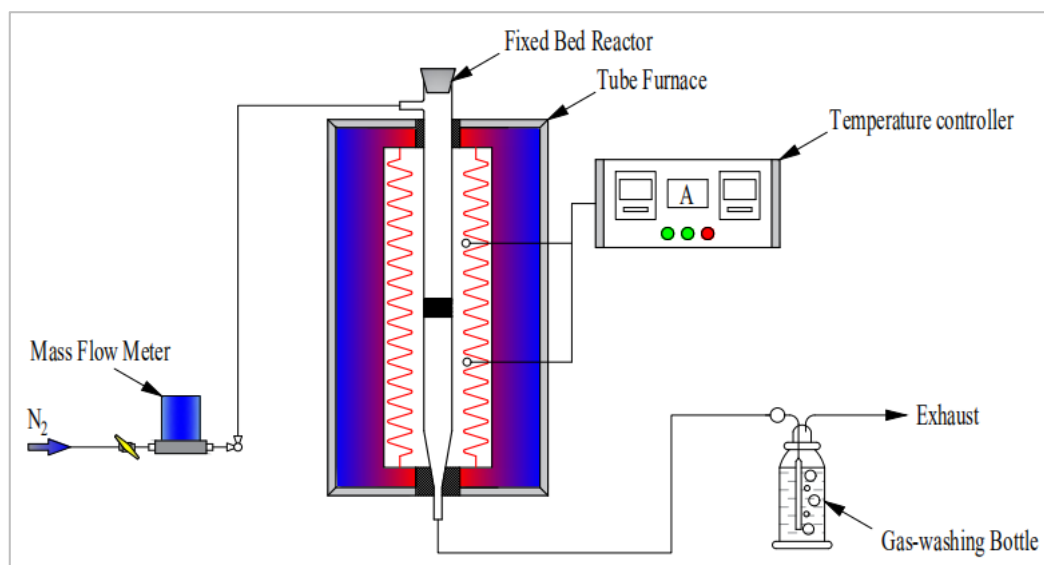
### **FTIR data:**

The FTIR spectrum of PAC-OP was similar to the two commercial ACs. The spectrum exhibited strong absorption band in the waveband of 3700–3200 cm<sup>-1</sup>, 1800–1300 cm<sup>-1</sup> and 1300–900 cm<sup>-1</sup>. The weak absorption peaks were seen at 2930, 2854, 2371, 2340, and 800 cm<sup>-1</sup>. Among them, the absorption band in the 3700–3200 cm<sup>-1</sup> can be ascribed to O–H stretching vibrations in hydroxyl, carboxylic, and

phenolic groups due to the existence of surface hydroxyl groups and chemisorbed water. The absorption band in the 2850–2950 and 1460–1350  $\text{cm}^{-1}$  can be assigned to stretching vibrations of aliphatic C–H groups. The absorption band in the 2400–2300  $\text{cm}^{-1}$  can be attributed to  $\text{CO}_2$  structure. The presence of bands at 1800–1300  $\text{cm}^{-1}$  can be attributed to the stretching vibrations of C=O in carboxylic acids, anhydrides, lactones, esters, carboxylic salts, carboxylo-carbonates, quinones, hydroxyquinones, C=C in carboxylic or pyridine-like structure. The presence of bands at 1300–900  $\text{cm}^{-1}$  range can be attributed to the stretching vibrations of C–O, C–O–C, OH in carboxylic acids, anhydrides, lactones, esters, ethers, phenols, epoxides and carboxyl-carbonates, which were also overlapped. The absorption band in the 900–700  $\text{cm}^{-1}$  can be attributed to the aromatic hydrocarbon groups, also, there is the C=N group located at 800  $\text{cm}^{-1}$  [S4-S7].

#### **PAC regeneration system and condition:**

Regeneration experiments of the samples after  $\text{SO}_2$  adsorption were completed in a thermal regeneration system, as shown in Fig. S2. This system is similar to the  $\text{SO}_2$  adsorption system, except that the regeneration gas is  $\text{N}_2$ , and the regeneration temperature is also different. The detailed regeneration process is as follows: PAC was heated in a flow of  $\text{N}_2$  (0.5 L/min), at a heating rate of 10  $^\circ\text{C}/\text{min}$ , and maintained at 450 $^\circ\text{C}$  for 1 hour and then cooled to room temperature while the  $\text{N}_2$  purge was continued.



**Fig. S2.** PAC regeneration system.

### References:

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