

Supplementary Materials for

“Forever chemicals”: a sticky environmental problem

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The method of calculating the half-life ($t_{1/2}$): The rate constant (k) needs to be estimated from the decrease in the pseudo-primary concentration over time, the natural logarithm of the reactant concentration is plotted in minutes, and the slope is given k (Equation 1). $t_{1/2}$ is calculated by Equation 2 [1].

$$\ln(C/C_0) = -kt \quad (1)$$

$$t_{1/2} = \ln 2/k \quad (2)$$

where C = concentration at time t , C_0 = initial concentration, k = rate constant, t = time, $t_{1/2}$ = half-life.

1. Schwarzenbach RP, Gschwend PM, and Imboden DM. Environmental organic chemistry. 3rd ed. Hoboken, N.J.: Wiley; c2017.

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